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## Synthesis and investigation of a hexyl substituted thieno-fused BODIPY derivative as a versatile near-IR fluorophore

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**Abstract:** The synthesis, photophysical and electrochemical properties of hexyl thiophene substituted thieno[*b*]-fused BODIPY structure are reported within this work. One such derivative, HTFBod, has been studied, which is one of the rare compounds that has a maximum absorbance wavelength greater than 750 nm among fused BODIPY compounds. In addition, it preserves its significant spectral properties such as high molar absorptivity and fluorescence quantum yield. Electrochemical characterizations indicated that the compound could be a successful candidate for an organic solar cell donor compound due to its low band gap, which is known to improve the short-circuit current ( $J_{sc}$ ) values. The compound could also harvest the near-IR portion of the solar irradiance where maximum solar photon flux exists. In addition, its low LUMO value of -4.13 eV contributes to its air stability. Hexyl substituents provide greater solubility which is required in solution-processable organic photovoltaics and further bathochromic shift when compared to similar compounds. Next, the synthesis was accomplished in five steps with higher yields compared to similar compounds in the literature. Heavy atom-free singlet oxygen production analysis has also been performed while it has been shown that the compound cannot produce singlet oxygen, while this property could be acquired *via* halogen substitution.

**Key words:** Dye, fluorophore, absorbance, boron dipyrromethene, bioimaging

### 1. Introduction

BODIPY compounds remain to be popular among various research groups for their versatile properties for the last few decades. These properties may be counted as large molar absorption coefficients, high fluorescence quantum yields, high stability, and derivatization from various sites.[1] Due to their rich chemistry, they can be employed as fluorescent chemosensors [2,3], fluorescent stains for bioimaging purposes [4], photosensitizers for organic photovoltaics [5], laser dyes [6], photodynamic therapy agents [7–9], active compounds in photocatalysis [10] and dyes for theranostic applications [11]. In recent years, there are various studies aiming to increase the number of near IR absorbing and emitting chemical dye compounds for some applications such as photodynamic therapy, noninvasive bioimaging agents, and organic photovoltaics. For example, in organic photovoltaics, it is known that solar irradiance is strong in the near IR region, therefore harvesting these is highly important to obtain highly efficient devices [12]. Primary BODIPY structures have been absorbing and emitting in the visible range around 500 nm. To shift these to higher wavelengths various strategies have been employed up to date. Extending the  $\pi$ -electron conjugation through the  $\alpha$ - and  $\beta$ - positions [13,14], aza-BODIPYs [15], and fused BODIPY[16] structures are among the most popular strategies. Fused BODIPY structures are classified as [*a*]-fused, [*b*]-fused and zig-zag-fused structures [17]. It has been observed that [*b*]-fused structures have more stabilization on HOMO and LUMO levels compared to the [*a*]-fused structures [17].

In the recent decade, there have been various reports [5,16,18–27] on [*b*]-fused BODIPY dyes for different applications which were recently reviewed [17]. Among these, heterocycle fused compounds attracted much attention, including thiophene and furan fused structures. Although these studies have enriched the properties of primary BODIPY dyes, more studies are required to improve the synthesis, photophysical properties and understand the structure-function relationship for their use in wider application areas. Due to these reasons, a novel hexyl substituted fluorophore based on [*b*]-thiophene fused BODIPY structure was developed and its photophysical, electrochemical properties, and singlet oxygen producing ability were investigated. Recently, our group have also investigated another fused BODIPY compound,

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namely, [a]-benzofused derivative [11]. In that research, we have successfully applied it on HeLa cells for phototoxic effect by comparing it with a novel distyryl derivative, besides exploiting other properties.

## 2. Materials and methods

### 2.1. General methods

All chemicals and solvents purchased from Aldrich, Acros, or TCI were used without further purification unless otherwise stated.  $^1\text{H-NMR}$  and  $^{13}\text{C-NMR}$  spectra were recorded using a Bruker DPX-400 in  $\text{CDCl}_3$  with TMS as an internal reference. Column chromatography of all products was performed using Merck Silica Gel 60 or Silicycle (particle size: 0.040–0.063 mm, 230–400 mesh ASTM). Reactions were monitored by thin layer chromatography using fluorescent coated aluminum sheets.

Absorption spectrometry in solution was performed using Agilent Cary 60 spectrophotometer. Steady state fluorescence measurements were conducted using an Agilent Eclipse spectrofluorometer. Solvents used for spectroscopy experiments were spectrophotometric grade. Fluorescence quantum yields were calculated by using the method in the literature [28] using the reference dye Zinc phthalocyanine excited at 650 nm in pyridine as the reference fluorophore ( $\Phi_f = 0.3$ ) [29].

Mass spectrometry (HRMS) measurements were performed using Agilent LC/MS-High Resolution Quadrupole Mass Time-of-Flight (Q-TOF) or Time-of-Flight (TOF).

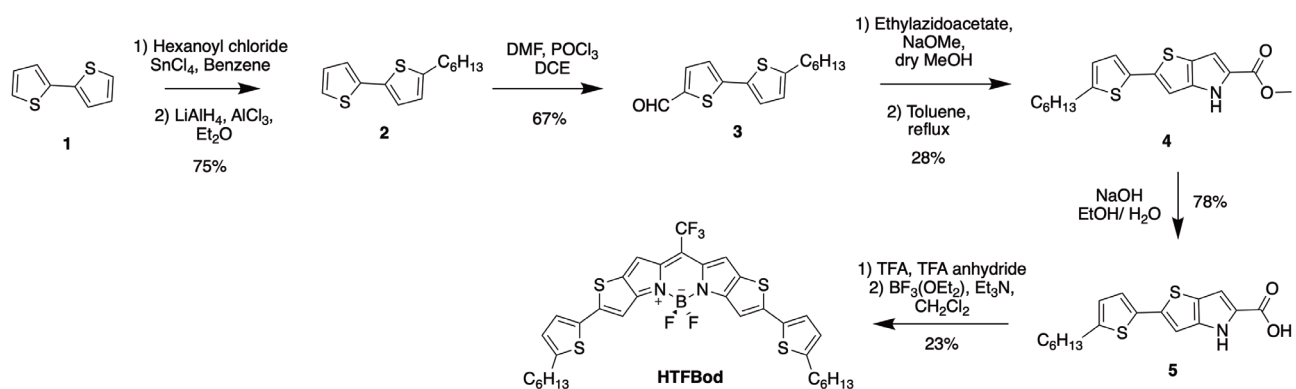
Cyclic voltammety was recorded on a Gamry Interface 1010E instrument in a single-component cell under a nitrogen atmosphere. A typical three-electrode configuration with a glassy carbon as working electrode, platinum wire as counter electrode and a saturated calomel electrode (SCE) as a reference electrode. Potentials were measured versus SCE and referenced to ferrocene as an internal standard ( $E^\circ(\text{Fc}/\text{Fc}^+) = -4.78$  eV to vacuum). The measurements were performed with a scan rate of  $100 \text{ mV}\cdot\text{s}^{-1}$  in degassed dichloromethane and the addition of tetra-n-butylammonium hexafluorophosphate (TBAPF,  $0.1 \text{ mol}\cdot\text{L}^{-1}$ ) as electrolyte. The compounds were used as  $1 \text{ mM}$  concentration.

Singlet oxygen measurements were measured according to the recently reported literature [11].

### 2.2. Synthesis of molecules

Molecules 2 and 3 were synthesized according to the literature [30] with the yields shown on Scheme.

**Compound 4:** Methanol (50 mL) was degassed with  $\text{N}_2$  gas for 15 min. Bithiophene aldehyde 3 (2 g, 6.85 mmol, 1 eq) was dissolved in degassed methanol (17 mL) and ethylazidoacetate (2.65 g, 20.55 mmol, 3 eq) was added. Separately, sodium methoxide (1.1 g, 20.55 mmol, 3 eq) was dissolved in degassed dry methanol (17 mL). The first flask was cooled to  $0^\circ\text{C}$  and a solution of sodium methoxide was added dropwise in half an hour and further stirred for another half an hour at this temperature. Then, at room temperature stirring was continued for 4-6 h and the reaction was stopped upon tracking the formation of the product with TLC (Hexane:Ethyl Acetate, 6:1). Then, sat.  $\text{NH}_4\text{Cl}$  was added to the reaction and a precipitate was formed. After the filtrate was removed, the precipitate was washed with ethyl acetate (50 mL) and extracted with water (50 mL). Then organic phase was dried with  $\text{Na}_2\text{SO}_4$  and the solvent was evaporated. Then, the compound was dissolved with toluene (50 mL) and degassed with  $\text{N}_2$  for 15 min, and then refluxed at  $120^\circ\text{C}$  for 1.5 h. The solvent was then evaporated and the pure compound was obtained after flash column chromatography purification (Hexane: ethyl acetate, 6:1). Dark yellow color in 28% yield.  $^1\text{H NMR}$  400 MHz,  $\text{CDCl}_3$ , ppm: (d) 9.02 (br s, 1H), 7.07 (m, 1H), 7.03 (m, 1H), 6.96 (m, 1H), 6.70 (m, 1H), 3.90 (s, 3H), 2.81 (m, 2H), 1.67 (m, 2H), 1.1–1.5 (m, 6H), 0.90 (m, 3H).  $^{13}\text{C NMR}$  (101 MHz,  $\text{CDCl}_3$ )  $\delta$  175.86, 146.43, 141.97, 135.59, 134.42, 130.00, 125.12, 124.05, 110.09, 108.13, 106.70, 51.96, 31.79, 30.44, 28.98, 22.80, 14.77, 14.32. (HRMS) $^-$ ; calculated for  $\text{C}_{18}\text{H}_{20}\text{NO}_2\text{S}_2$  (M-H) $^-$  as 346.0941, found 346.0956.



**Scheme.** Synthesis of target molecule HTFBod.

**Compound 5:** Compound 4 (241 mg, 0.693 mmol, 1 eq) and NaOH (580 mg, 14.55 mmol, 21 eq) were dissolved in ethanol (12 mL) and water (6 mL). The reaction mixture was refluxed for 1 h and then cooled to room temperature. By the addition of 6 M HCl, the pH was adjusted to 3 and then the formed precipitate was filtrated. The precipitate was dissolved with ethyl acetate and after drying with  $\text{Na}_2\text{SO}_4$ , the solvent was evaporated. No further purification was performed. Off-white solid with 78% yield.  $^1\text{H}$  NMR 400 MHz,  $\text{DMSO-d}_6$ , ppm: (d) 11.6 (br s, 1H), 7.07 (m, 1H), 6.97 (s, 1H), 6.75 (m, 2H), 2.75 (m, 2H), 1.60 (m, 2H), 1.4–1.1 (m, 6H), 0.85 (m, 3H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{DMSO}$ )  $\delta$  163.05, 145.40, 141.81, 139.06, 135.88, 126.00, 123.94, 122.30, 108.04, 106.90, 31.73, 31.65, 30.10, 28.79, 22.73, 14.60. (HRMS) $^-$ ; for  $(\text{M-H})^- \text{C}_{17}\text{H}_{18}\text{NO}_2\text{S}_2^-$  calculated as 332.0784, found as 332.0798.

**HTFBod:** Compound 5 (50 mg, 0.15 mmol) was dissolved with trifluoroacetic acid and stirred at 40 °C for half an hour. Subsequently, trifluoroacetic anhydride (40  $\mu\text{L}$ ) was added and heated to 80 °C and stirred for 45 min at this temperature. Then, after cooling to room temperature, sat.  $\text{NaHCO}_3$  in ice-water was slowly added attentively (10 mL). Then, the formed dark green precipitates were filtrated, and the precipitates were dissolved with DCM and dried with  $\text{Na}_2\text{SO}_4$ . And the solvent was evaporated. Then, the compound was dissolved in toluene (5 mL) and triethylamine (0.2 mL) and boron trifluoride diethyl ether (0.2 mL) were added. The reaction flask was stirred at room temperature for 1 h and the solvent was evaporated and purified by column chromatography (DCM: Hexane, 1:1). Dark green color with 23% yield.  $^1\text{H}$  NMR 400 MHz,  $\text{CDCl}_3$ , ppm:  $\delta$  7.28 (d,  $J = 3.7$  Hz, 2H), 7.20 (s, 2H), 7.16 (s, 2H), 6.79 (d,  $J = 3.8$  Hz, 2H), 2.85 (t,  $J = 7.6$  Hz, 5H), 1.71 (p,  $J = 7.4$  Hz, 5H), 1.46–1.30 (m, 15H), 0.95–0.84 (m, 7H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ )  $\delta$  160.58, 155.31, 151.94, 138.84, 135.41, 134.21, 128.07, 126.36, 124.48, 117.80, 108.18, 31.74, 31.61, 30.77, 29.93, 28.96, 22.78, 14.29. (HRMS) $^-$ ; for  $(\text{M-H})^- \text{C}_{34}\text{H}_{33}\text{BF}_5\text{N}_2\text{S}_4^-$  calculated as 703.1545, found as 703.15386.

### 3. Results and discussion

#### 3.1. Synthetic studies

For the synthesis of HTFBod, five steps of synthesis were employed. In the beginning, it was started with commercially available 2,2'-bithiophene. By using hexanoyl chloride and tin (IV) chloride in benzene, a hexanoyl unit was attached. Later, it was reduced to the alkyl form by using  $\text{LiAlH}_4$  and  $\text{AlCl}_3$  in diethyl ether to form compound 2 in 75% yield. Next, it was formylated by using the Vilsmeier-Haack reaction conditions in 67%. These two steps were accomplished according to the literature [30]. The key chemical ethylazidoacetate was synthesized in our laboratory by using ethyl bromoacetate and sodium azide *via* substitution reaction by refluxing the reaction mixture for 4 h in acetone. Then, compound 3 was reacted with ethylazidoacetate in basic conditions in dry methanol through Hemetsberger-Knittel indolization reaction in two steps. The first intermediate is prone to decomposition, therefore, cyclization reaction in toluene was accomplished to obtain a more stable compound 4 in 28% yield. Similar thiophene-substituted fused-thiophene BODIPY derivative has also been synthesized by You group [19,20]. However, they followed a different route: they first synthesized a bromo thieno-pyrrole-fused BODIPY and attached the edge thiophene groups in the later stage through Suzuki coupling and they reported the above-stated indolization reaction yield as 5% in their most recent study to form the bromo substituted thieno-pyrrole. Then, with this strategy, it was proposed a better yield of synthesis of this key intermediate by starting with alkyl-bithiophene rather than attaching the thiophene subsequently through Suzuki coupling.

After obtaining compound 4, it was employed basic hydrolysis to obtain compound 5 in decent yield and without purification. In the next step, compound 5 was reacted with trifluoroacetic acid and trifluoroacetic anhydride. The former chemical is used for decarboxylation reaction, while the latter has been used to form meso-trifluoromethyl substitution. After the formation of the dipyrromethene derivative boron trifluoride diethyl ether complex and triethylamine were added in dichloromethane to form the HTFBod in 23% yield.  $^1\text{H}$ ,  $^{13}\text{C}$  NMR and high-resolution mass spectroscopy characterizations were employed to confirm the structures.

#### 3.2. Photophysical characterizations

To characterize this near IR dye, several studies were employed. First, absorbance spectroscopy studies were investigated in various organic solutions with 5  $\mu\text{M}$  concentration (Figure 1). Thanks to the bis-hexyl units, the compound is well-soluble in organic solvents. The maximum absorption wavelength of HTFBod in chloroform is 764 nm. This compound can absorb between 650 and 810 nm range. The molar absorptivity of this compound has been found as 202,000  $\text{M}^{-1}\text{cm}^{-1}$  in chloroform (Table 1). The maximum fluorescence in the same solvent was obtained at 788 nm. Both the absorbance and emission spectra of this compound are quite sharp with small Stokes shift. A similar observation was also observed by the Suzuki group with Keio Fluor dyes for the [b]-furan-fused BODIPY compounds [25]. The full width at half maximum (fwhm) values is similar between these two studies. Then, the fluorescence quantum yield was calculated as 0.11 by using reference dye zinc phthalocyanine. Therefore, since the HTFBod has moderate fluorescence, it could be utilized for bioimaging/imaging purposes. Absorbance and fluorescence characteristics of this dye were also accomplished in four

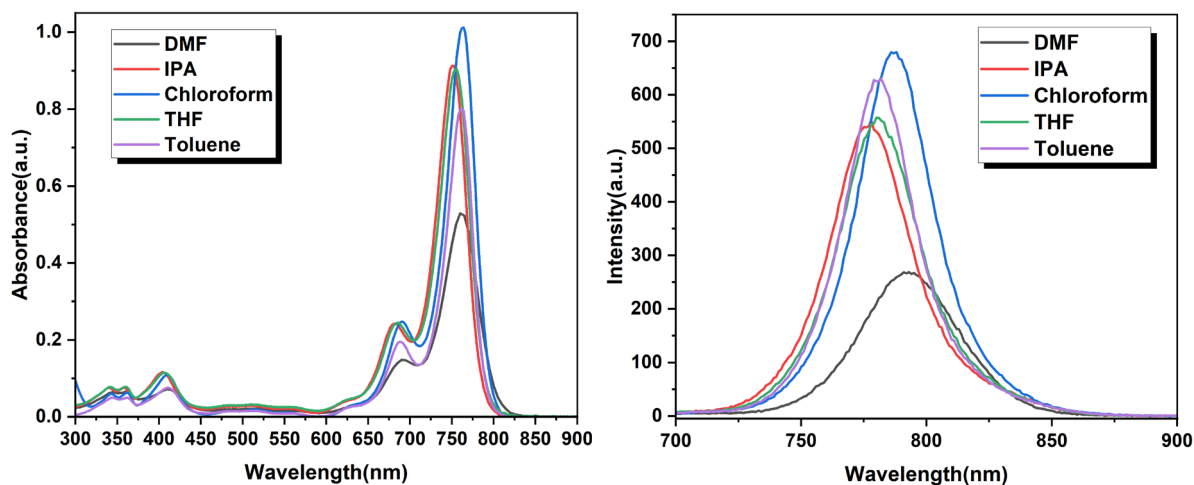


Figure 1. Absorbance (left) and fluorescence (right) spectra of HTFBod in different solvents (5  $\mu\text{M}$ ).

Table 1. Spectroscopic data for HTFBod in various solvents.

Compound	$\lambda_{\text{max}}$ (abs) nm	$\lambda_{\text{max}}$ (ems) nm	$\epsilon_{\text{max}}$ ( $\text{M}^{-1}\text{cm}^{-1}$ )	Dn ( $\text{cm}^{-1}$ )	Solvent	$\Phi_f^a$
HTFBod	764	788	202,000	399	$\text{CHCl}_3$	0.11
	752	778	182,000	444	IPA	
	756	780	180,000	407	THF	
	762	779	160,000	286	Toluene	
	760	793	106,000	548	DMF	

<sup>a</sup>Fluorescence quantum yield was measured in chloroform and calculated by using Zinc phthalocyanine excited at 650 nm in pyridine as the reference fluorophore ( $\Phi_f = 0.3$ ) [29]

other solvents: dimethylformamide, iso-propyl alcohol, toluene, and tetrahydrofuran (Table 1). It was observed that the compound has similar characteristics in terms of absorbing-emission wavelength range in nonpolar and polar solvents: the absorbance maxima are in 750–760 nm range and emission in 780–790 nm range. The fluorescence has not been disappeared in polar media.

There are few reports on compounds which have maximum absorbance in the region where HTFBod absorbs among fused BODIPY structures. Kubota, Matsui synthesized [b]-thiophene fused BODIPY compounds with dimethylaniline attached to the thiophene unit to obtain highly redshifted compounds with 783 and 813 nm absorption maxima due to the strong electron donating effect of the amino unit [18]. In comparison with another related study by the You group, where they investigated thiophene substituted thiophene-fused BODIPY structure [20], it was observed 30 nm bathochromic shift with HTFBod due to the electron donating effect of the alkyl (hexyl) units. The photophysical properties of HTFBod set it in one of the rare fluorophores, that have absorption maxima greater than 750 nm among the fused-BODIPY structures. As far as our knowledge, there is no such compound which has maximum absorption ( $\lambda_{\text{max}}$ ) larger than 750 nm and as high molar absorptivity and fluorescence quantum yield than HTFBod [17,31].

Previous studies have shown that attaching electron withdrawing units to the meso position and electron donating units to the 2- and 3- positions of the BODIPY core is a quite useful method for shifting the absorbance wavelength to the red end. Therefore, the employment of electron deficient trifluoromethyl unit and electron-rich bithiophene unit attachment seems to be an effective way of obtaining near IR dyes. In comparison to the similar compound by D'Souza and You, it has been observed that substitution with thiophene rather than phenylene unit has *ca.* 60 nm bathochromic (709 nm vs. 764 nm) [19].

### 3.3. Electrochemical characterization

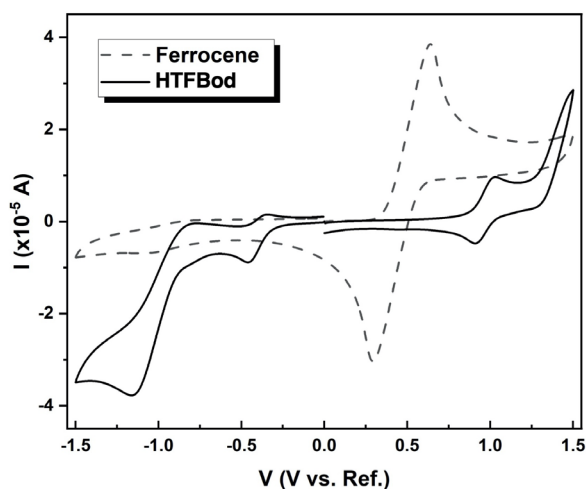
Synthesized near-IR dye HTFBod was examined by cyclic voltammetry (CV) to obtain the oxidation and reduction potentials and the corresponding frontier orbital energy levels (HOMO and LUMO energy levels) and the bandgap. The

measurements were accomplished in dichloromethane by using a saturated calomel electrode (SCE) as a reference electrode (The ferrocene/ferrocenium (Fc/Fc<sup>+</sup>) couple was also measured separately as a redox standard) [32]. The voltammograms of HTFBod and ferrocene have been shown in Figure 2.

For the HTFBod, scans were performed between +1.5 and -1.5 V, and one reversible oxidation peak was observed in this region at about +0.89 V, showing a stable Bodipy cation based on the earlier studies [30,33]. Two irreversible reduction waves are also observable with onsets of -0.32 V and -0.84 V.

By using the half-wave oxidation and reduction potentials of the BODIPY compound, the HOMO and LUMO energies have been calculated with the potential of Fc/Fc<sup>+</sup> as a reference energy of  $E_{(\text{HOMO})} = -4.78$  eV [32,34]. For the HTFBod, HOMO is estimated at -5.34 eV, while the LUMO is -4.13 eV (Table 2). The corresponding bandgap from the cyclic voltammetry measurement was calculated as 1.21 eV. The optical bandgap was also estimated from the onset of absorbance spectrum by the equation given in the literature as  $E_g = 1242/\lambda_{\text{onset}}$  and calculated accordingly as 1.53 eV [35]. The bandgap estimated from the absorbance spectrum is a bit larger than the value obtained from the CV spectrum. Similar observations have also been recorded in other studies [30,32]. Thin-film absorbance data may give closer values.

Decreasing the LUMO energy level of dyes/fluorophores is an efficient way of designing air-stable compounds and is desired for many applications such as organic photovoltaics (OPV) [36]. The presence of a very low LUMO energy level of HTFBod of -4.13 eV is significant in this aspect. Furthermore, the energy levels of the frontier orbitals indicate that HTFBod can be a suitable candidate for organic solar cell donor material for low band gap compounds. It is known that low band gap compounds produce increased short-circuit current ( $J_{sc}$ ) which increases the efficiency [37]. However, frequently used organic solar cell acceptor material PC<sub>61</sub>BM and PC<sub>71</sub>BM have LUMO energy levels around -4.0 eV (The LUMO of the PC<sub>61</sub>BM was measured as -4.09 eV) [30] and this value should be lower than the donor compound's LUMO energy level, which is -4.13 eV. Therefore, compatible with this data other nonfullerene acceptor compounds would suit. Otherwise, some modifications may also be accomplished on HTFBod, such as insertion of phenyl groups instead of trifluoromethyl unit in the meso position. Our observations on the BODIPY core suggest that electron withdrawing groups on this position stabilize the LUMO energy level.



**Figure 2.** Scans of HTFBod and Ferrocene measured in  $\text{CH}_2\text{Cl}_2/\text{Bu}_4\text{PF}_6$  (0.1 M) versus SCE (scan rate  $100 \text{ mVs}^{-1}$ ).

**Table 2.** Electrochemical data of the HTFBod, and calculated HOMO-LUMO energy values.

Molecule	$E_{\text{ox onset}}(\text{V})^a$	HOMO (eV) <sup>b</sup>	$E_{\text{red onset}}(\text{V})^a$	LUMO (eV) <sup>b</sup>	$E_g$ (eV) <sup>c</sup>	$E_g$ (eV) <sup>d</sup>
<b>HTFBod</b>	+0.89	-5.34	-0.32, -0.84	-4.13	1.21	1.53

Half-wave potentials of the first oxidation and first reduction, measured in  $\text{CH}_2\text{Cl}_2/\text{Bu}_4\text{PF}_6$  (0.1 M) versus SCE, scan rate  $100 \text{ mVs}^{-1}$ , with Fc/Fc<sup>+</sup> as standard was measured separately and  $E_{\text{ox}}^{1/2}(\text{Fc})$  was found as +0.33 V. [b]  $E_{\text{HOMO}}(\text{Fc}) = -4.78$  eV,  $E_{\text{HOMO}} = -4.78 + (E_{\text{ox}}^{1/2}(\text{Fc}) - (E_{\text{ox}}^{1/2}))$ ,  $E_{\text{LUMO}} = -4.78 + (E_{\text{ox}}^{1/2}(\text{Fc}) - (E_{\text{red}}^{1/2}))$ . [c] HOMO-LUMO difference as obtained from CV. [d] based on spectroscopic data in solution.

As a successful application of fused BODIPY structures to OPV, Ma and Leo group have synthesized [b]-furan fused BODIPY structures and fabricated organic solar cell devices using these compounds as donor compounds by vacuum deposition method and achieved one of the highest efficiencies within BODIPY-based compounds in tandem devices (10%) [5].

### 3.4. Singlet oxygen formation studies

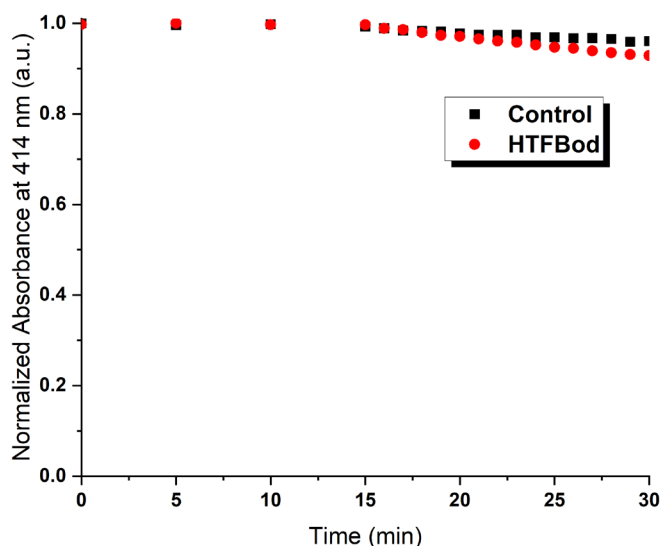
Singlet oxygen generation without using heavy halogen atoms has been extensively studied over the last 10 years [4,7,38,39]. Fused BODIPY structures have also been studied in photodynamic therapy methods for both halogenated [21] and halogen-free versions. Interestingly, in the halogen-free versions, You [19] and Nguyen, Yoon [23] groups have published several thiophene fused structures and have observed singlet oxygen formation in some of the derivatives. They have noted that both the meso substituent and the unit substituted directly to the thiophene play a role. It was observed that for  $^1\text{O}_2$  production ability, when the trifluoromethyl group attached to the meso position, the substituent attached to the thiophene unit should not be electron rich.

In this case, it was studied the singlet oxygen formation studies with HTFBod to test if it produces singlet oxygen or not. To perform experiment by chemical methods diphenylisobenzofuran (DPBF) as a singlet oxygen scavenger compound was used in organic solvent isopropanol. By using HTFBod and the LED irradiation (700 nm) in an air bubbled solution, the change in absorbance spectrum of the scavenger compound was measured (see Figure 3). However, has not been observed any significant decrease in the absorbance, thereby it was concluded that the compound is not an efficient  $^1\text{O}_2$  producer.

The result is also coherent by the observations of the previous study [19]. Attachment of electron-rich hexyl-thiophene structure to the fused thiophene BODIPY unit is expected not to produce  $^1\text{O}_2$ . They have also estimated that the energy level difference between the HOMO and the LUMO should be at least larger than 1.53 eV. However, in our case, this value was estimated as 1.21 eV. Therefore, producing  $^1\text{O}_2$  with nonhalogenated near-IR absorbing compounds becomes more ineffective as the absorption wavelength increases. However, it is possible to produce by halogen substitution [22].

## 4. Conclusion

In this study, the synthesis and investigation of a novel derivative of [b]-thiophene fused BODIPY structure were employed. Fused BODIPY structures could be very versatile building blocks for various application areas. This type of structure successfully employed in organic photovoltaics (both dye-sensitized and organic solar cells), photodynamic therapy, and bioimaging studies. Upon investigating various methods to obtain fused BODIPY structures using Hemetsberger-Knittel indolization reaction, it was performed five steps of synthesis starting from the commercially available and cheap 2,2'-bithiophene structure. It has been shown that by using inexpensive chemicals it is possible to reach thiophene substituted thiophene-fused BODIPY structures in a more efficient way compared to the similar compounds in the literature. With



**Figure 3.** Time-dependent decrease in normalized absorbance (414 nm) of DPBF as a result of oxidation by HTFBod (5  $\mu\text{M}$ ) in aerated IPA upon irradiation with 700 nm LED array.

an absorbance maximum at 764 nm and emission maxima at 788 nm (in CHCl<sub>3</sub>) it is one of the most redshifted fused compounds. In addition, it has an acceptable fluorescence quantum yield of 0.11, which is almost zero in most of the other near-IR fused compounds ( $J_{\max} > 750$  nm) [19,20]. Therefore, HTFBod is one of the rare fused compounds that could be used in this wavelength range. Investigation of the electrochemical studies revealed that the low band gap of the HTFBod can be suitable for organic photovoltaics applications since low band gap compounds have higher  $J_{sc}$  values. Furthermore, near IR absorbing compounds are desired in OPV studies because of the fact that the maximum solar photon flux is located between 600 and 800 nm [40]. However, more suitable nonfullerene acceptors should be employed due to the low LUMO energy level of the compound. Alternatively, in order to increase the LUMO level of the HTFBod, it is proposed to use an aromatic unit instead of trifluoromethyl unit in the meso position of the BODIPY. Our studies based on developing fused structures for applications in various areas are currently under study in our research group.

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### References

1. Kue C, Ng S, Voon S, Kamkaew A, Chung L et al. Recent strategies to improve boron dipyrromethene (BODIPY) for photodynamic cancer therapy: an updated review. *Photochemical & Photobiological Sciences* 2018; 17: 1691-1708.
2. Bozdemir O, Guliyev R, Buyukcakir O, Selcuk S, Kolemen S et al. Selective manipulation of ICT and PET processes in styryl-Bodipy derivatives: applications in molecular logic and fluorescence sensing of metal ions. *Journal of the American Chemical Society* 2010; 132: 8029-8036.
3. Bozdemir OA, Sozmen F, Buyukcakir O, Guliyev R, Cakmak Y et al. Reaction-based sensing of fluoride ions using built-in triggers for intramolecular charge transfer and photoinduced electron transfer. *Organic Letters* 2010; 12: 1400-1403.
4. Nguyen V, Yim Y, Kim S, Ryu B, Swamy K et al. Molecular design of highly efficient heavy-atom-free triplet Bodipy derivatives for photodynamic therapy and bioimaging. *Angewandte Chemie-International Edition* 2020; 59: 8957-8962.
5. Li T, Meyer T, Ma Z, Benduhn J, Korner C et al. Small molecule near-infrared boron dipyrromethene donors for organic tandem solar cells. *Journal of the American Chemical Society* 2017; 139: 13636-13639.
6. Zhang D, Martin V, Garcia-Moreno I, Costela A, Perez-Ojeda M et al. Development of excellent long-wavelength BODIPY laser dyes with a strategy that combines extending pi-conjugation and tuning ICT effect. *Physical Chemistry Chemical Physics* 2011; 13: 13026-13033.
7. Cakmak Y, Kolemen S, Duman S, Dede Y, Dolen Y et al. Designing excited states: theory-guided access to efficient photosensitizers for photodynamic action. *Angewandte Chemie-International Edition* 2011; 50: 11937-11941.
8. Türkoğlu G, Kayadibi KG, Yurt MNZ, Pirencioglu SN, Erbas-Cakmak S A therapeutic keypad lock decoded in drug resistant cancer cells. *Chemical Science* 2021; 12: 9754-9758.
9. Arslan B, Bilici K, Demirci G, Almammadov T, Khan M et al. A leucine aminopeptidase activatable photosensitizer for cancer cell selective photodynamic therapy action. *Dyes and Pigments* 2021; 195: 109735.
10. Luo G, Fang K, Wu J, Mo J. Photocatalytic water reduction from a noble-metal-free molecular dyad based on a thienyl-expanded BODIPY photosensitizer. *Chemical Communications* 2015; 51: 12361-12364.
11. Can B, Cakmak Y. Exploration of two different strategies in near ir absorbing boron dipyrromethene derivatives for photodynamic and bioimaging purposes. *ChemistrySelect* 2021; 6: 8855-8860.
12. Koppe M, Egelhaaf H, Dennler G, Scharber M, Brabec C et al. Near IR sensitization of organic bulk heterojunction solar cells: towards optimization of the spectral response of organic solar cells. *Advanced Functional Materials* 2010; 20: 338-346.
13. Cakmak Y, Akkaya EU Phenylethynyl-bodipy oligomers: bright dyes and fluorescent building blocks. *Organic Letters* 2009; 11: 85-88.
14. Zhu S, Zhang J, Vegesna G, Tiwari A, Luo F et al. Controlled knoevenagel reactions of methyl groups of 1,3,5,7-tetramethyl bodipy dyes for unique bodipy dyes. *RSC Advances* 2012; 2: 404-407.
15. Adarsh N, Avirah R, Ramaiah D. Tuning photosensitized singlet oxygen generation efficiency of novel aza-bodipy dyes. *Organic Letters* 2010; 12: 5720-5723.
16. Sun Y, Yu X, Yang J, Gai L, Tian J et al. NIR halogenated thieno[3, 2-b]thiophene fused BODIPYs with photodynamic therapy properties in HeLa cells. *Spectrochimica Acta Part a-Molecular and Biomolecular Spectroscopy* 2021; 246.



17. Wang J, Boens N, Jiao L, Hao E. Aromatic [b]-fused BODIPY dyes as promising near-infrared dyes. *Organic & Biomolecular Chemistry* 2020; 18: 4135-4156.
18. Kubota Y, Kimura K, Jin J, Manseki K, Funabiki K et al. Synthesis of near-infrared absorbing and fluorescing thiophene-fused BODIPY dyes with strong electron-donating groups and their application in dye-sensitized solar cells. *New Journal of Chemistry* 2019; 43: 1156-1165.
19. Watley R, Awuah S, Bio M, Cantu R, Gobeze H et al. Dual functioning thieno-pyrrole fused bodipy dyes for nir optical imaging and photodynamic therapy: singlet oxygen generation without heavy halogen atom assistance. *Chemistry-an Asian Journal* 2015; 10: 1335-1343.
20. Awuah S, Das S, D'Souza F, You Y. Thieno-pyrrole-fused bodipy intermediate as a platform to multifunctional NIR agents. *Chemistry-an Asian Journal* 2013; 8: 3123-3132.
21. Yang Y, Guo Q, Chen H, Zhou Z, Guo Z et al. Thienopyrrole-expanded BODIPY as a potential NIR photosensitizer for photodynamic therapy. *Chemical Communications* 2013; 49: 3940-3942.
22. Awuah S, Polreis J, Biradar V, You Y. Singlet oxygen generation by novel nir bodipy dyes. *Organic Letters* 2011; 13: 3884-3887.
23. Qi S, Kwon N, Yim Y, Nguyen V, Yoon J. Fine-tuning the electronic structure of heavy-atom-free BODIPY photosensitizers for fluorescence imaging and mitochondria-targeted photodynamic therapy. *Chemical Science* 2020; 11: 6479-6484.
24. Umezawa K, Matsui A, Nakamura Y, Citterio D, Suzuki K. Bright, color-tunable fluorescent dyes in the vis/NIR region: establishment of new "tailor-made" multicolor fluorophores based on borondipyrromethene. *Chemistry-a European Journal* 2009; 15: 1096-1106.
25. Umezawa K, Nakamura Y, Makino H, Citterio D, Suzuki K. Bright, color-tunable fluorescent dyes in the visible-near-infrared region. *Journal of the American Chemical Society* 2008; 130: 1550-+.
26. Wang J, Li J, Chen N, Wu Y, Hao E et al. Synthesis, structure and properties of thiophene-fused BODIPYs and azaBODIPYs as near-infrared agents. *New Journal of Chemistry* 2016; 40: 5966-5975.
27. Wu Y, Cheng C, Jiao L, Yu C, Wang S et al. Beta-thiophene-fused BF<sub>2</sub>-azadipyrromethenes as near-infrared dyes. *Organic Letters* 2014; 16: 748-751.
28. Mirloup A, Leclerc N, Rihn S, Bura T, Bechara R et al. A deep-purple-grey thiophene-benzothiadiazole-thiophene BODIPY dye for solution-processed solar cells. *New Journal of Chemistry* 2014; 38: 3644-3653.
29. Seybold PG, Gouterman M. Porphyrins : XIII: Fluorescence spectra and quantum yields. *Journal of Molecular Spectroscopy* 1969; 31: 1-13.
30. Bura T, Leclerc N, Fall S, Leveque P, Heiser T et al. High-performance solution-processed solar cells and ambipolar behavior in organic field-effect transistors with thienyl-BODIPY scaffoldings. *Journal of the American Chemical Society* 2012; 134: 17404-17407.
31. Jean-Gerard L, Vasseur W, Scherninski F, Andrioletti B. Recent advances in the synthesis of [a]-benzo-fused BODIPY fluorophores. *Chemical Communications* 2018; 54: 12914-12929.
32. Gresser R, Hummert M, Hartmann H, Leo K, Riede M. Synthesis and characterization of near-infrared absorbing benzannulated aza-bodipy dyes. *Chemistry-a European Journal* 2011; 17: 2939-2947.
33. Ziessel R, Bonardi L, Retailleau P, Ulrich G. Isocyanate-, isothiocyanate-, urea-, and thiourea-substituted boron dipyrromethene dyes as fluorescent probes. *Journal of Organic Chemistry* 2006; 71: 3093-3102.
34. Connelly N, Geiger W. Chemical redox agents for organometallic chemistry. *Chemical Reviews* 1996; 96: 877-910.
35. Leonat L, Sbarcea G, Branzoi IV. Cyclic voltammetry for energy levels estimation of organic materials. *UPB Scientific Bulletin Series B* 2013; 75: 111-118.
36. Wakamiya A, Murakami T, Yamaguchi S. Benzene-fused BODIPY and fully-fused BODIPY dimer: impacts of the ring-fusing at the b bond in the BODIPY skeleton. *Chemical Science* 2013; 4: 1002-1007.
37. Kang H, Kim G, Kim J, Kwon S, Kim H et al. Bulk-heterojunction organic solar cells: five core technologies for their commercialization. *Advanced Materials* 2016; 28: 7821-7861.
38. Duman S, Cakmak Y, Kolemen S, Akkaya E, Dede Y. Heavy atom free singlet oxygen generation: doubly substituted configurations dominate S-1 states of bis-Bodipys. *Journal of Organic Chemistry* 2012; 77: 4516-4527.
39. Ji S, Ge J, Escudero D, Wang Z, Zhao J et al. Molecular structure-intersystem crossing relationship of heavy-atom-free bodipy triplet photosensitizers. *Journal of Organic Chemistry* 2015; 80: 5958-5963.
40. Mishra A, Bauerle P. Small molecule organic semiconductors on the move: promises for future solar energy technology. *Angewandte Chemie-International Edition* 2012; 51: 2020-2067.

## SUPPLEMENTARY MATERIAL

### **Synthesis and Investigation of a Hexyl Substituted Thieno-Fused BODIPY Derivative as a Versatile Near IR**

#### **Fluorophore**

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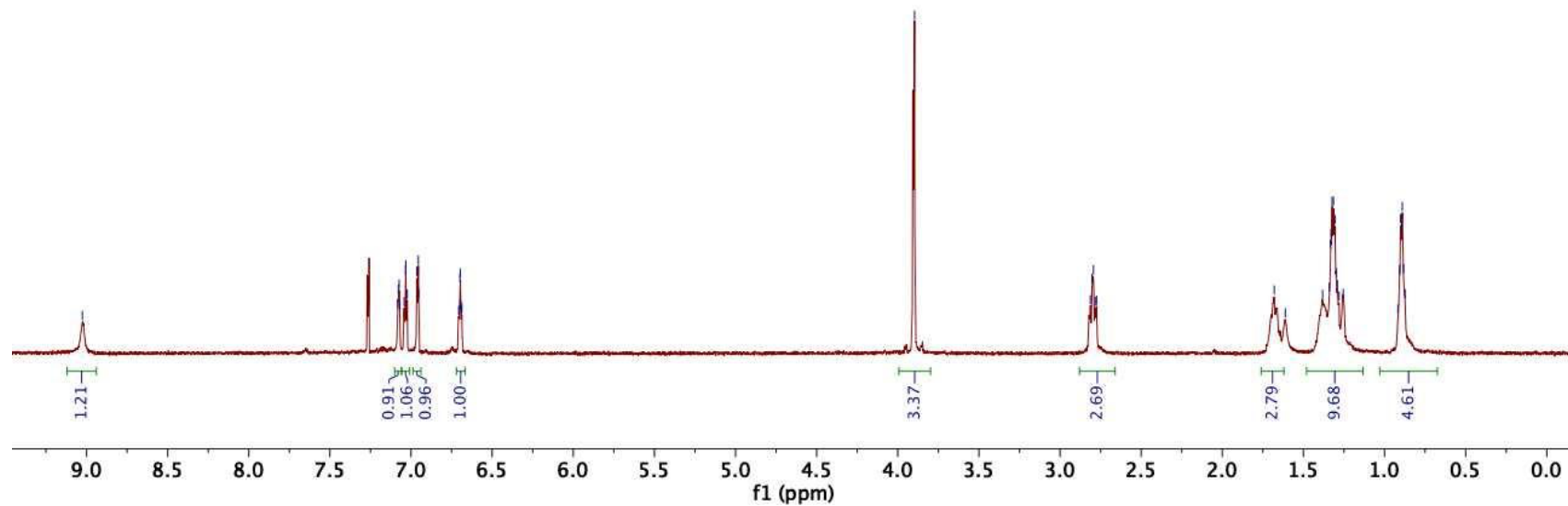
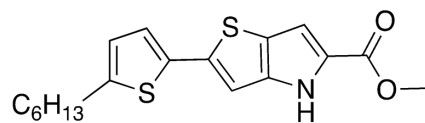
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— 3609.08  
 2832.14  
 2830.33  
 2828.26  
 2826.42  
 2816.05  
 2812.68  
 2811.75  
 2808.32  
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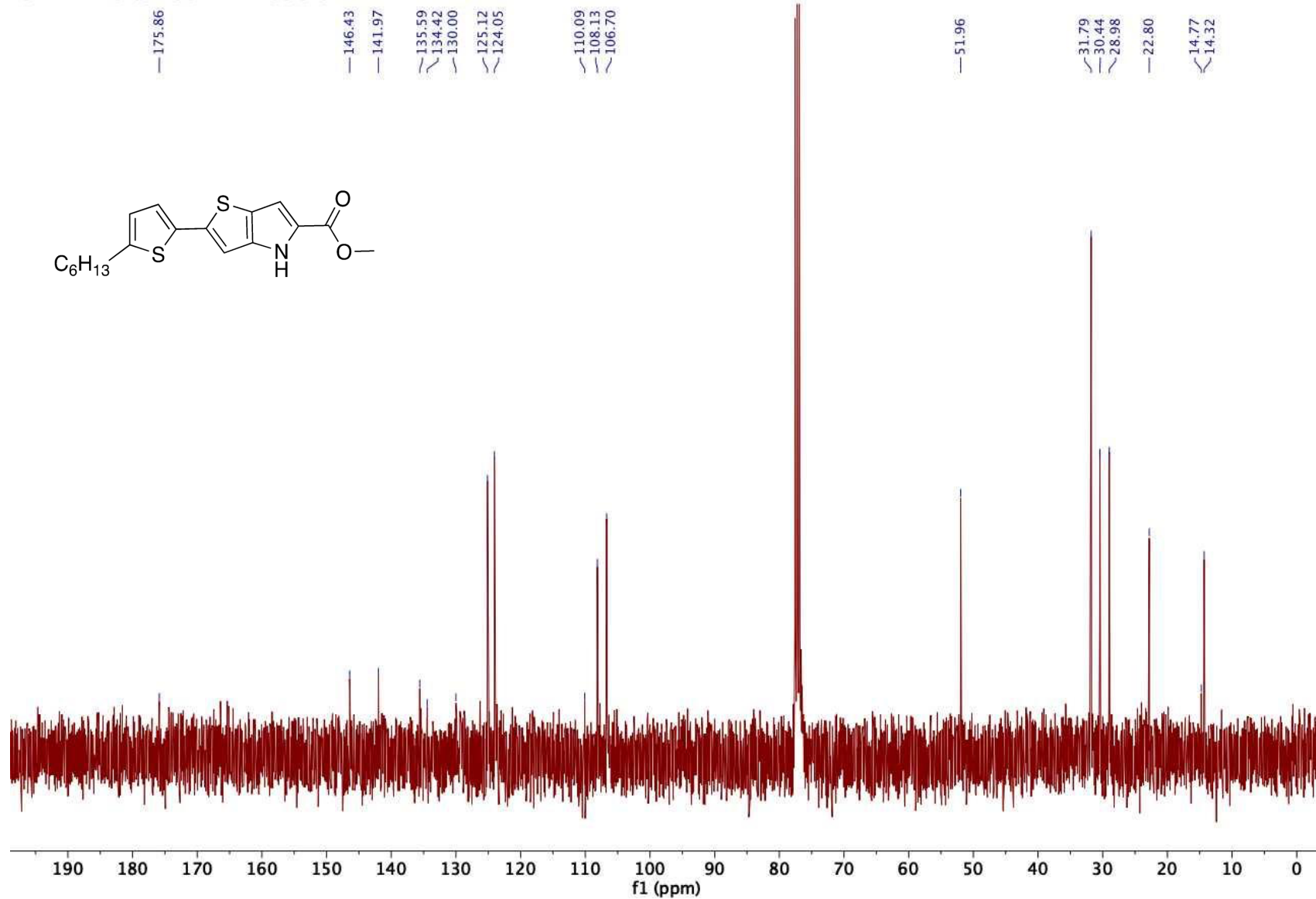
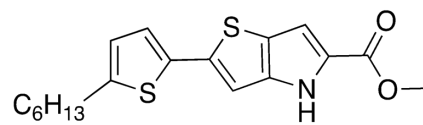
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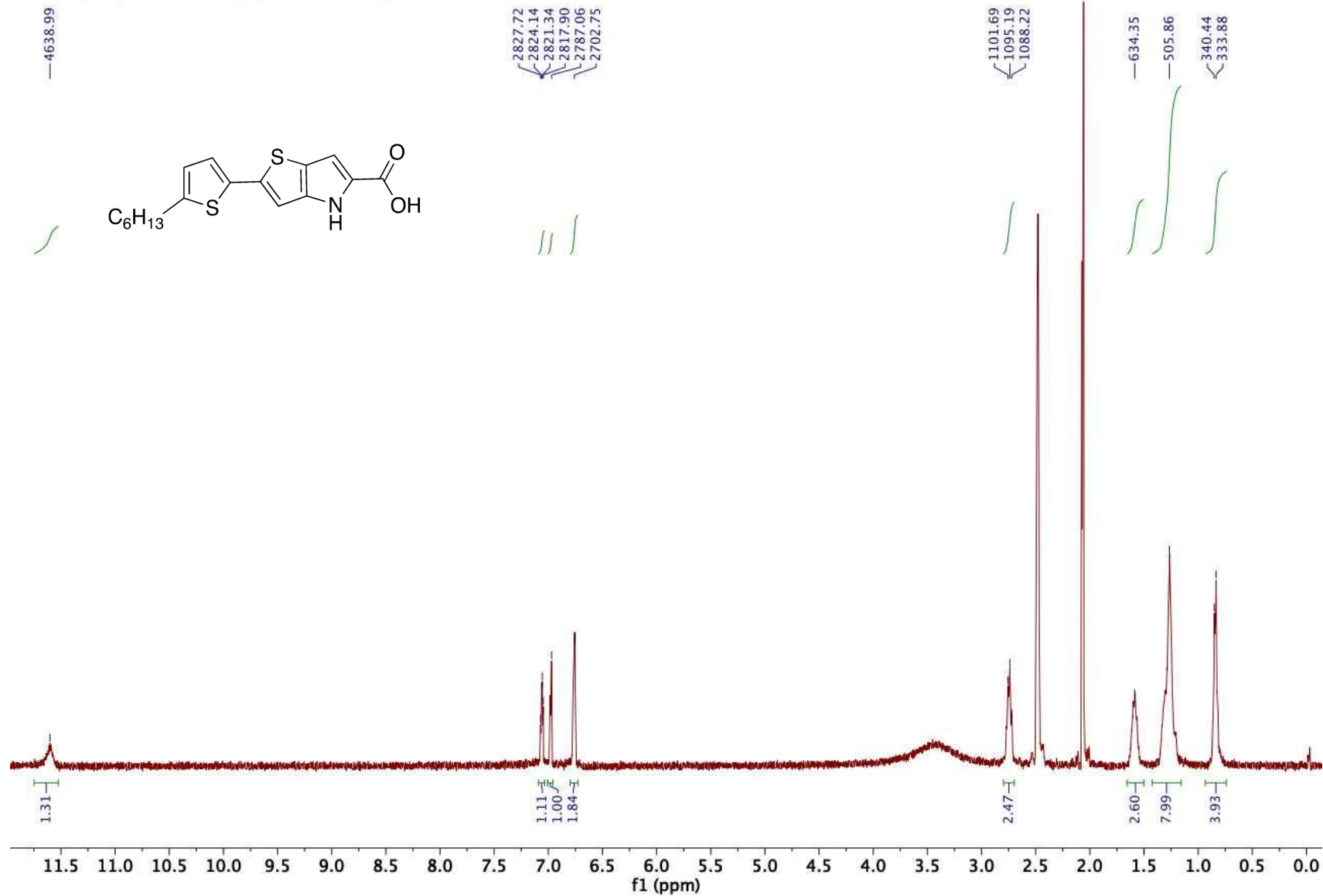
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 501.64  
 363.03  
 360.29  
 357.66  
 355.98  
 353.10  
 340.24



YCAKMAK-YC103-13CNMR — 13C OBSERVE — —



mt162-r3 — STANDARD PROTON PARAMETERS —



BC24-13CNMR — 13C OBSERVE — —

163.05

145.40

141.81

139.06

135.88

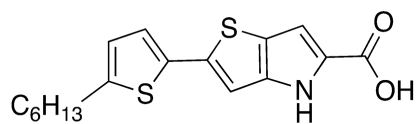
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108.04

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31.73

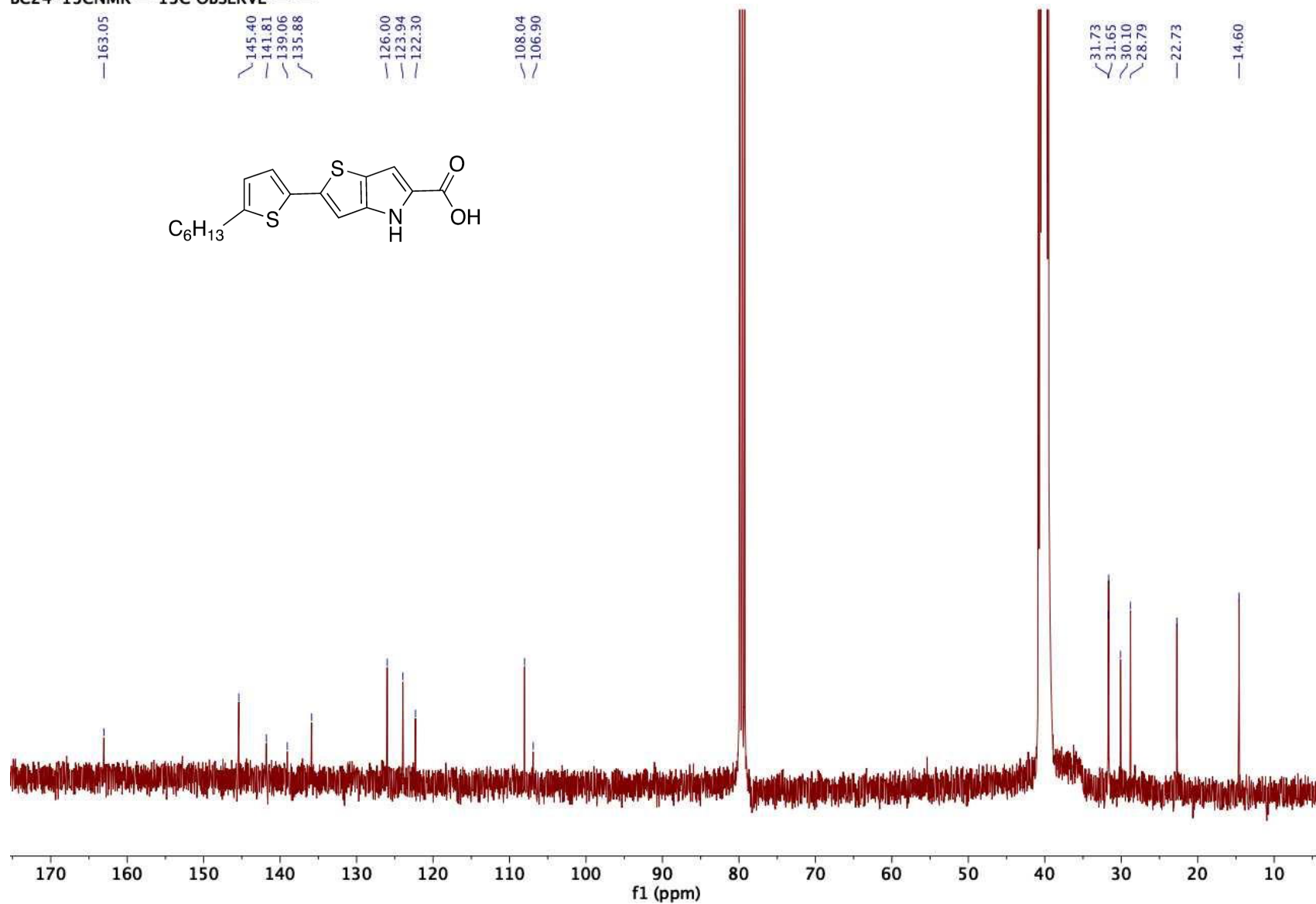
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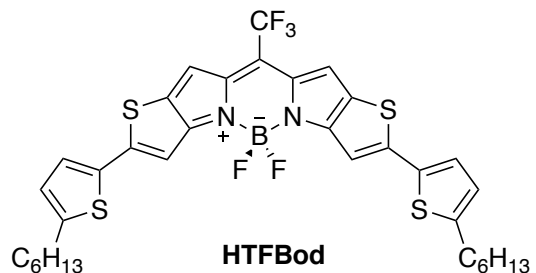
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YCAKMAK-D4-1HNMR — STANDARD 1H OBSERVE — —

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2879.83  
2877.50  
2862.48  
2718.59  
2714.82

1147.11  
1139.55  
1131.94  
697.48  
690.37  
683.13  
675.57  
667.88  
665.79  
564.28  
556.95  
550.05  
544.18  
535.95  
532.48  
528.90  
524.83  
521.43  
369.26  
362.57  
355.81



C (s)  
7.20

B (s)  
7.16

A (d)  
6.79

D (d)  
7.28

E (t)  
2.85

F (p)  
1.71

G (m)  
1.33

H (m)  
0.90

